**T04D03 - Lewis Dot Structures & NAS Rule**

Draw the following Lewis Dot structures for molecules or ions that may or may not follow the octet rule.

Use the formula N-A = S and if it has an ***even number of shared electron*** then

* it follows the octet rule unless ……..

it is group IIA (Be) or IIIA (B) or gets more than 8 electrons around it such as in SCl6 ,etc.

If it has an ***odd number of shared electrons*** then

* it does not follow the octet rule and the central atom will receive the odd electron

a. CS2 b. SO3-2

c. NO2 d. NO2-1

e. HCN f. AsF5

g. BeCl2 h. N2F2 (FNNF)

i. SO3 j. NO